

Honors Chemistry: Solutions 6 Problem Set- Solution Stoichiometry

1) How many grams of silver chloride are produced when 40.0 mL of 0.25 M silver nitrate is reacted with excess sodium chloride? The other product is sodium nitrate.

2). In this reaction, 40.0 mL of a 0.100 M sodium bicarbonate reacts with an excess of hydrochloric acid. How many liters of carbon dioxide are produced at STP? The other products are water and sodium chloride.

4) A total of 4.50 grams of lead(II) chromate is produced from the reaction of 100 mL of sodium chromate and 100 mL lead(II) nitrate. What is the initial molarity of the lead(II) nitrate?

5) A 50.0 mL sample of 0.75 M potassium iodide reacts with 50.0 mL of 0.75 M lead(II) nitrate. How many grams of precipitate are formed?