

Honors Chemistry: Atomic Theory 12 Problem Set – Light and Planck's Equation

1. Complete the table below. You will need to use the electromagnetic spectrum and Planck's light equations. The spectrum can be found under the "additional resources" section.

Energy (Joules)	Wavelength- λ (meters)	Frequency - ν (s^{-1})	Color of Light/type of electromagnetic radiation
6.3×10^{-19} J			
	2.4×10^{-7} m		
		$100 s^{-1}$	
1.5×10^{-14} J			
	5.6×10^{-10} km		
		$2.2 \times 10^{13} s^{-1}$	
	525 nm		

