

## Acid/Base Podcast #7: Redox Practice Problems; SC7a

Indicate the oxidation states of each of the elements in the following chemical species:

- 1) NaOH \_\_\_\_\_
- 2) HBr \_\_\_\_\_
- 3) VF<sub>5</sub> \_\_\_\_\_
- 4) Na<sub>2</sub>SO<sub>4</sub> \_\_\_\_\_
- 5) NO<sub>2</sub><sup>-1</sup> \_\_\_\_\_
- 6) SiOF<sub>2</sub> \_\_\_\_\_
- 7) Fe<sub>3</sub>O<sub>4</sub> \_\_\_\_\_
- 8) Zn(OH)<sub>4</sub><sup>-2</sup> \_\_\_\_\_

Determine which of the following reactions are redox reactions. For those that are, identify the element which has been oxidized, and the one that has been reduced. Next, write the half reactions for all redox reactions.

